

CHEMISTRY

(Academic Sessions 2009 – 2011, 2010 – 2012 and 2011 – 2013)

PAPER – I (Objective Type) 212-(INTER PART – I)

Time Allowed : 20 Minutes

GROUP – I

PAPER CODE = 2485

Maximum Marks : 17

Note : Four possible answers A, B, C and D to each question are given. The choice which you think is correct, fill that circle in front of that question with Marker or Pen ink. Cutting or filling two or more circles will result in zero mark in that question. Write the letter A, B, C or D in the column (write correct option) against each question also. If there is a contradiction in the bubble and hand written answer, bubble option will be considered correct.

1-1	The partial pressure of oxygen in the lungs is : (A) 760 torr (B) 670 torr (C) 159 torr (D) 116 torr
2	The term pH was introduced by : (A) Henderson (B) Sorenson (C) Goldstein (D) J.J. Thomson
3	The existence of an element in more than one crystalline form is called : (A) Allotropy (B) Isotopy (C) Isomorphism (D) Polymorphism
4	The number of bonds in nitrogen molecule is : (A) One sigma and one pi (B) One sigma and two pi (C) Three sigma only (D) Two sigma and one pi
5	The mass of one mole of electrons is : (A) 1.008 mg (B) 0.55 mg (C) 0.184 mg (D) 1.673 mg
6	Splitting of spectral lines when atoms are subjected to strong electric field is called : (A) Zeeman effect (B) Stark effect (C) Compton effect (D) Photoelectric effect
7	In zero order reaction, the rate is independent of : (A) Temperature of reaction (B) Concentration of reactants (C) Concentration of products (D) Concentration of reactants and products
8	Molarity of pure water is : (A) 1 (B) 18 (C) 55.5 (D) 6
9	Reduction always takes place at : (A) Anode (B) Cathode (C) Both electrodes (D) Salt bridge
10	The number of moles of CO ₂ which contain 8.0 g of oxygen is : (A) 0.25 (B) 0.50 (C) 1.0 (D) 1.50
11	The nature of the positive rays depend on : (A) The nature of the cathode (B) The nature of the anode (C) The nature of the residual gas (D) The nature of the discharge tube
12	The oxidation number of chromium in K ₂ Cr ₂ O ₇ is : (A) 14 (B) 12 (C) 6 (D) 13
13	Which of the following molecule has zero dipole moment : (A) NH ₃ (B) CHCl ₃ (C) H ₂ O (D) BF ₃
14	If an endothermic reaction is allowed to take place very rapidly in the air, the temperature of the surrounding air : (A) Remain constant (B) Increases (C) Decreases (D) Remain unchanged
15	Pressure remaining constant, at which temperature the volume of a gas will become twice of what it is at 0 °C : (A) 546 °C (B) 200 °C (C) 546 K (D) 273 K
16	Acetone and chloroform are soluble in each other due to : (A) Intermolecular hydrogen bonding (B) Ion-dipole interaction (C) Instantaneous dipole (D) Covalent bonding
17	The drying agent used in a desiccator is : (A) AgCl (B) NH ₄ Cl (C) CaCl ₂ (D) AlCl ₃